



ZINCI SULFATIS LIQUIDUM PERORALUM PAEDIATRICUM PAEDIATRIC ZINC SULFATE ORAL SOLUTION

Final text for addition to *The International Pharmacopoeia*

This monograph was adopted at the Forty-second WHO Expert Committee on Specifications for Pharmaceutical Preparations in October 2007 for addition to the 4th edition of The International Pharmacopoeia.

[Note from the Secretariat: The term "paediatric" has been included in the title of this monograph since this oral solution is included in the 15th WHO Model list of essential medicines (EML) (March 2007) under "medicines for diarrhoea in children" (section 17.5.2) and in the same section of the 1st WHO Model list of essential medicines for children (October 2007).]

Category. Adjunct to oral rehydration salts in (prevention and) treatment of dehydration due to diarrhoea.

Storage. Paediatric zinc sulfate oral solution should be kept in a well-closed container.

Labelling. The designation of the container of Paediatric zinc sulfate oral solution should indicate the quantity in terms of the equivalent amount of elemental zinc.

Additional information. Strength in the current WHO Model list of essential medicines and in the WHO Model list of essential medicines for children: 10 mg of zinc (as zinc sulfate monohydrate or zinc sulfate heptahydrate) per 5 ml. Additional strength available: 20 mg of zinc (as zinc sulfate monohydrate or zinc sulfate heptahydrate) per 5 ml.

[Note from the Secretariat. This additional strength is given since, although not in the EMLs, it is included in the WHO/UNICEF Guidelines on Production of Zinc tablets and Zinc Oral Solution.]

Requirements

Complies with the monograph for "Liquid Preparations for Oral Use".

Definition. Paediatric zinc sulfate oral solution is a solution of Zinc Sulfate as the monohydrate or heptahydrate in a suitable flavoured vehicle. It contains not less than 90.0% and not more than 110.0% of the amount of zinc stated on the label.

Manufacture. The formulation of the oral solution and the manufacturing process are designed and controlled so as to ensure that the metallic taste of the zinc salt is adequately masked.

Identity tests

- A. To 5 ml add 0.2 ml of sodium hydroxide (400 g/l) TS. A white precipitate is formed. Add a further 2 ml of sodium hydroxide (400 g/l) TS. The precipitate dissolves. Add 10 ml of ammonium chloride (100 g/l) TS. The solution remains clear. Add 0.1 ml of sodium sulfide TS. A flocculent white precipitate is formed.
- B. Five ml yields reaction A described under 2.1 General identification tests as characteristic of sulfates.

pH value. (1.13) pH of the oral solution: 2.5 - 4.5.

Relative density. (1.3) Relative density of the oral solution: 1.18 - 1.24.

Assay. To a quantity of the oral solution equivalent to about 10 mg of zinc, accurately measured, add 50 ml of water R and 5 ml of ammonia buffer TS and titrate with disodium edetate (0.01 mol/l) VS, using about 50 mg of Mordant Black 11 indicator mixture R as indicator, until the solution turns from violet to blue. Each ml of disodium edetate (0.01 mol/l) VS is equivalent to 0.6539 mg of zinc.
